

SC 5051
WASSCE SC 2018
CHEMISTRY 1
Objective Test
1 hour

1

Name:.....

Index Number:.....

THE WEST AFRICAN EXAMINATIONS COUNCIL

West African Senior School Certificate Examination for School Candidates, 2018

SC 2018

CHEMISTRY 1

1 hour

OBJECTIVE TEST

Do not open this booklet until you are told to do so. While you are waiting, write your name and index number in the spaces provided at the top right-hand corner of this booklet and thereafter, read the following instructions carefully.

1. Use **HB** pencil throughout.
2. If you have got a blank answer sheet, complete its top section as follows.
 - (a) In the space marked *Name*, write in capital letters your **surname** followed by your **other names**.
 - (b) In the spaces marked *Examination, Year, Subject* and *Paper*, write ‘WASSCE SC’, ‘2018’, ‘CHEMISTRY’ and ‘1’ respectively.
 - (c) In the box marked *Index Number*, write your **index number** vertically in the spaces on the left-hand side. There are numbered spaces in line with each digit. Shade carefully the space with the same number as each digit.
 - (d) In the box marked *Paper Code*, write the digits **505113** in the spaces on the left-hand side. Shade the corresponding numbered spaces in the same way as for your index number.
 - (e) In the box marked *Sex*, shade the space marked **M** if you are **male**, or **F** if you are **female**.
3. If you have got a pre-printed answer sheet, check that the details are correctly printed, as described in 2 above. In the boxes marked *Index Number, Paper Code* and *Sex*, **reshade** each of the shaded spaces.
4. An example is given below. This is for a **male** candidate whose name is **Chinedu Oladapo DIKKO**, whose index number is **4251102068** and who is offering **CHEMISTRY 1**.

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PRINT IN BLOCK LETTERS

Name: DIKKO CHINEDU OLADAPO Examination: WASSCE SC Year: 2018
Surname Other Names

Subject: CHEMISTRY Paper: 1

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PAPER CODE	
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SEX
Indicate your sex by shading the space marked M (for Male) or F (for Female) in this box: M F <input type="checkbox"/> <input type="checkbox"/>

For Supervisors only.
If candidate is absent shade this space:

- INSTRUCTIONS TO CANDIDATES**
1. Use grade **HB** pencil throughout.
 2. Answer each question by choosing one letter and shading it like this: [A] [B] [C]
 3. Erase completely any answer(s) you wish to change.
 4. Leave extra spaces blank if the answer spaces provided are more than you need.
 5. Do not make any markings across the heavy black marks at the right-hand edge of your answer sheet.

Answer **all** Questions

Each question is followed by **four** options lettered A to D. Find out the **correct** options for each question and shade in **pencil** on your answer sheet, the answer space which bears the same letter as the option you Chosen. Give only **one** answer to **each** question. An example is given below

Which of the following elements reacts with water?

- A. Carbon
- B. Iodine
- C. Sodium
- D. Sulphur

The correct answer is Sodium, which is lettered C and therefore answer space C would be shaded.

[A] [B] [C] [D]

Think carefully before you shade the answer spaces; erase completely any answer you wish to change.

Do **all** rough work on this questions paper.

Now answer the following questions.

1. Which of the following raw materials is used in the plastic industry?
 - A. Ethene
 - B. Methane
 - C. Sulphur
 - D. Hydrogen

2. Which of the following organic compounds can undergo both addition and substitution reactions?
 - A. Petane
 - B. Benzene
 - C. Propane
 - D. Hexane

3. Which of the following equations represents a redox reaction?
 - A. $\text{AgNO}_{3(\text{aq})} + \text{KCl}_{(\text{aq})} \rightarrow \text{AgCl}_{(\text{s})} + \text{KNO}_{3(\text{aq})}$
 - B. $\text{HNO}_{3(\text{aq})} + \text{NaOH}_{(\text{aq})} \rightarrow \text{NaNO}_{3(\text{aq})} + \text{H}_2\text{O}_{(\text{l})}$
 - C. $\text{CaCO}_{3(\text{s})} \rightarrow \text{CaO}_{(\text{s})} + \text{CO}_2(\text{g})$
 - D. $2\text{H}_2\text{S}(\text{g}) + \text{SO}_2(\text{g}) \rightarrow 2\text{H}_2\text{O}_{(\text{l})} + 3\text{S}_{(\text{g})}$

4. The process of extraction of iron from its ore is
 - A. decomposition.
 - B. oxidation.
 - C. reduction.
 - D. sublimation.

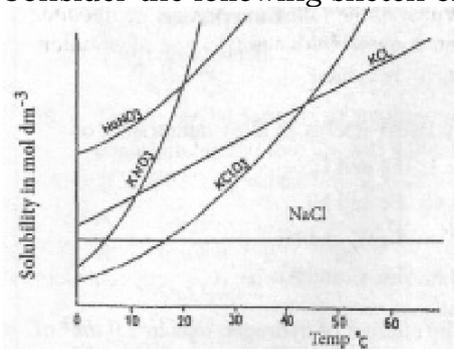
5. What is the solubility of a salt if 0.4 g of it is obtained on evaporating 200 cm³ of its saturated solution to dryness?
 - A. 0.08 gdm⁻³
 - B. 2.00 gdm⁻³
 - C. 8.00 gdm⁻³
 - D. 80.00 gdm⁻³

6. An acidic salt has
 - A. double anions in its aqueous solution.
 - B. a single cation in its aqueous solution.
 - C. hydrogen ions in its aqueous solution.
 - D. hydrogen atoms in its aqueous solution.

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7. A reaction is endothermic if the
- reaction vessel feels cool during the reaction.
 - enthalpy change is negative.
 - bond forming energy exceeds bond breaking energy.
 - heat of formation of reactants exceeds heat of formation of products.
8. In which of the following compounds does hydrogen form ionic compounds?
- CH₄
 - HCl
 - NH₃
 - NaH
9. Consider the following reaction equation:
 $\text{Br}_2 + 2\text{KI} \rightarrow 2\text{KBr} + \text{I}_2$
Bromine is acting as
- an oxidizing agent.
 - a reducing agent.
 - an acid.
 - a base.
10. An organic compound has the empirical formula CH₂. If its molar mass is 42 gmol⁻¹ what is its molecular formula?
[H = 1.0, C = 12.0]
- C₂H₄
 - C₃H₄
 - C₃H₆
 - C₄ H₈
11. Ethene is produced from ethanol by
- decomposition.
 - hydrolysis.
 - ozonolysis.
 - dehydration.
12. Consider the following equilibrium reaction:
 $2\text{AB}_{2(\text{g})} + \text{B}_{2(\text{g})} \rightleftharpoons 2\text{AB}_{3(\text{g})}$ $\Delta H = -X\text{kJmol}^{-1}$
The backward reaction will be favored by
- a decrease in pressure.
 - an increase in pressure.
 - a decrease in temperature.
 - an introduction of a positive catalyst.
13. What is the mass of solute in 500 cm³ of 0.005 mol dm⁻³ H₂SO₄?
[H = 1.0, O = 16.0, S = 32.0]
- 0.490 g
 - 0.049 g
 - 0.245 g
 - 0.0245 g
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14. Pure water can be made to boil at a temperature lower than 100 °C by
- reducing its quantity.
 - decreasing the external pressure.
 - distilling it.
 - increasing the external pressure.

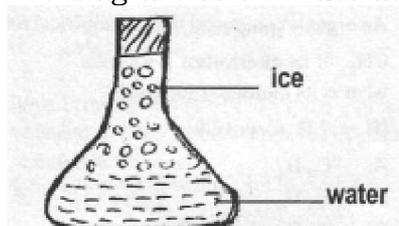
15. Consider the following sketch of the solubility curve of some substances.



At what temperature does the solubility of KNO₃ equal that of NaNO₃?

- A. 0°C
B. 20°C
C. 30°C
D. 40°C
16. When a salt is added to its saturated solution, the salt
A. dissolves and the solution becomes super saturated.
B. dissolves and the solution becomes unsaturated.
C. precipitates and the solution remains unchanged.
D. dissolves and crystals are formed.
17. When substance X was added to a solution of bromine water, the solution became colorless. X is likely to be
A. propane.
B. propanoic acid.
C. propyne.
D. propanol.
18. The preferential discharge of ions during electrolysis is influenced by the
A. mechanism of electrolysis.
B. electrolytic reactions.
C. nature of the electrode.
D. type of electrolytic cell.
19. The valence electrons of ${}_{12}\text{Mg}$ are in the
A. 3s orbital.
B. 2p_x orbital.
C. 2s orbital.
D. 1s orbital.
20. Stainless Steel is an alloy comprising of
A. Fe and C.
B. Fe and Ni.
C. Fe, C and Ni.
D. Fe, C and Al.
21. The number of hydrogen ions in 1.0 dm³ of 0.02 mol dm⁻³ tetraoxosulphate(VI) acid is [N_A = 6.02 x 10²³]
A. 1.2 x 10²²
B. 1.2 x 10²³.
C. 2.4 x 10²².
D. 2.4 x 10²³.

22. The most suitable substance for putting out petrol fire is
- water.
 - carbon(IV)oxide.
 - fire blanket.
 - sand.
23. The following factors would contribute to environmental pollution except
- production of ammonia.
 - manufacture of cement.
 - photosynthesis.
 - combustion.
24. The position of equilibrium in a reversible reaction is affected by
- particle size of the reactants.
 - vigorous stirring of the reaction mixture.
 - presence of a catalyst.
 - change in concentration of the reactants.
25. The diagram below illustrates a conical flask containing water and ice.



- Which of the following statements about the diagram is correct?
- The water is at a lower temperature than the ice
 - Energy is absorbed when the ice changes to water
 - Energy is released when the ice changes to water
 - The water molecules vibrate about a fixed point
26. Which of the following statements best explains the differences between a gas and a vapor?
- Unlike gases, vapors are liquids at room temperature
 - Unlike gases, vapor can easily be condensed into liquids
 - Unlike gases, vapour is readily converted into solids
 - Vapours are generally denser than gases
27. Consider the following reaction equation:
 $2\text{HCl} + \text{Ca}(\text{OH})_2 \rightarrow \text{CaCl}_2 + \text{H}_2\text{O}$
 What is the volume of 0.1 mol dm^{-3} HCl that would completely neutralize 25 cm^3 of 0.3 mol dm^{-3} $\text{Ca}(\text{OH})_2$?
- 150 cm^3
 - 75 cm^3
 - 30 cm^3
 - 25 cm^3
28. Cu and HNO_3 are not suitable for preparing hydrogen gas because of their
- reactivity and oxidation respectively.
 - conductivity and corrosiveness respectively.
 - melting point and reduction respectively.
 - electro negativity and solubility respectively.
29. Which of the following formulae cannot be an empirical formula?
- CH
 - CH_2
 - P_2O_5
 - N_2O_4

30. One of the criteria for confirming the purity of benzene is to determine its
- A. heat capacity.
 - B. boiling point.
 - C. mass.
 - D. colour.

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